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► **To cite this version:**

Marcel Hollenstein. Nucleic acid enzymes based on functionalized nucleosides. *Current Opinion in Chemical Biology*, 2019, 52, pp.93-101. 10.1016/j.cbpa.2019.06.007 . pasteur-02185135

HAL Id: pasteur-02185135

<https://pasteur.hal.science/pasteur-02185135>

Submitted on 21 Dec 2021

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HEAT TRANSFER IN SALT CAVERNS

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Abstract

Several aspects of the thermodynamic behavior of salt caverns containing brine, oil, natural gas, air or hydrogen are discussed. It is shown that natural convection always appears in a liquid-filled cavern. This is not true in the case of gas caverns. Gas type, brine temperature at the cavern bottom and water vapor content are influential. Equations are obtained for the thermodynamic behavior of the fluids stored in a cavern during injection-withdrawal. Thermal evolution of a liquid-filled cavern at rest is slow; rapid injection and withdrawal are adiabatic and lead to small temperature changes. Heat capacity is much smaller in a gas cavern, temperature evolutions are much faster, and heat transfer from the rock mass must be taken into account. Evolutions can be described through a relatively simple equation. A simple model describing pressure and temperature in the wellbore during gas withdrawal is obtained.

Key words: Salt caverns; Gas storage caverns; CAES; Thermodynamics of salt caverns; Water vapor in a salt cavern

Highlights

- Conditions leading to natural convection in a cavern
- Inversion of temperature gradient in a gas cavern
- Equations for evolution of stored fluid temperature during injection-withdrawal
- Adiabatic and non-adiabatic behavior of a liquid-storage cavern

- Thermodynamic behavior of a gas-storage cavern
- A closed-form solution for gas flow in a wellbore

Introduction

Salt caverns are deep cavities that are created by solution mining in both bedded and domal salt formations. They are connected to the ground level through a cased and cemented well. Strings are set in the well to allow injection or withdrawal of fluids into or from the cavern. These caverns range in volume from $V_c = 5000 \text{ m}^3$ to $5,000,000 \text{ m}^3$ and in depth from $H = 200 \text{ m}$ to $H = 2000 \text{ m}$. When solution mining is completed, the cavern can be filled with crude oil, LPG, natural gas, hydrogen, compressed air, etc. During operation, fluids are injected or withdrawn from the cavern, and heat is exchanged with the rock mass, leading to pressure and temperature changes in the cavern. An abundant literature — mentioned in the course of the paper — was dedicated to various aspects of these thermodynamic effects in specific sites. The objective of this paper is to describe some of these effects and to prove that their main features can be captured by relatively simple models.

1. Natural convection in a salt cavern

In this section, the theoretical conditions leading to the onset of natural convection in a cavern are discussed.

1.1. Fluid pressure in a cavern

In a gas-storage cavern, gas pressure (P) at casing-shoe depth (H) varies between a maximum and a minimum pressure, typically $P_{\min} = \gamma_{\min} H$, $P_{\max} = \gamma_{\max} H$, where $\gamma_{\min} = 0.6 \times 10^{-2} \text{ MPa/m}$, and $\gamma_{\max} = 1.8 \times 10^{-2} \text{ MPa/m}$ are the minimum and maximum “gradients”. Maximum gas pressure is smaller than geostatic pressure, $P_{\infty} = \gamma_R H$, $\gamma_R = 2.2 \times 10^{-2} \text{ MPa/m}$. where γ_R is the average volumetric weight of the rock formation. In natural gas caverns, gas used to be injected in summer and withdrawn in winter; at a $H = 1300\text{-m}$ depth, for instance, pressure varied from 8 MPa to 22 MPa, typically. Prompted by the needs of energy traders, cycles tend to be more frequent. In compressed air storage,

pressure is cycled every day between 5 MPa and 7 MPa, typically. A gas-storage cavern contains gas and brine (hence, water vapor) as, during the initial gas fill (debrining) some brine, trapped into the sump at the cavern bottom, cannot be removed from the cavern. This brine, which is cold (see Section 4.2), plays a significant role in cavern thermodynamics.

In liquefied or liquid (or supercritical) products storage, a central tube is left in the wellbore, and brine is injected in the central tube when hydrocarbon is withdrawn from the annulus between the cemented casing and the central tube (and vice versa): most often, cavern pressure is halmostatic, or $P_h = \gamma_b H$, $\gamma_b = 1.2 \times 10^{-2} \text{ MPa/m}$ is brine volumetric weight. However, cavern pressure is higher than halmostatic during tightness tests and lower during workovers. In the following, cavern pressure (P_c) is the fluid pressure computed at casing-shoe depth.

1.2. Temperature in a cavern

1.2.1. Brine-filled cavern

Figure 1 displays the temperature profile along a vertical axis in the cavern and access well of a brine cavern at Etrez, France, operated by Storengy (¹Brouard, 1998). For this cavern, $V_c = 8\,000 \text{ m}^3$, and the cavern's top and bottom depths are 920 m and 960 m, respectively. On the left hand-side, a slope discontinuity can be observed at a 700-m depth. Its origin is in the contrast between the thermal conductivities of salt and marls. Heat flux through conduction is continuous through any discontinuity, $q = K_R dT_R / dz = K_{marl} dT_{marl} / dz$. For rock salt, $K_R = 6 \text{ W/m/}^\circ\text{C}$; for marls, $K_{marl} = 3 \text{ W/m/}^\circ\text{C}$. The thermal gradient dT / dz is divided by 2 at the interface between salt and marls. Another slope discontinuity, originating in convection, can be observed at shallow depth (a few dozen meters below ground level). The test is performed in February. At night, the wellhead is cold. Cold brine flows down in the wellbore to be replaced by warmer brine from below which, in turn, cools when in contact with the wellhead. The natural geothermal gradient is $G_R^\infty = dT_R^\infty / dz = 1.6 \times 10^{-2} \text{ }^\circ\text{C/m}$ in the salt formation. The thermal conductivity of the brine is $K_b = 0.6 \text{ W / m / }^\circ\text{C}$: if heat transfer in the cavern brine were through conduction, one would expect that the temperature gradient be larger in the brine mass than in the rock mass. In fact, the temperature gradient in the cavern is as small as $G_c = 3.4 \times 10^{-3} \text{ }^\circ\text{C/m}$, clear evidence of the effects of natural convection. At the cavern bottom, the brine is warmer (hence, lighter) than it is at the cavern top. Due to

buoyancy effects, brine rises; it cools along its upward path to the cavern top and ultimately returns to the cavern bottom, leading to the creation of convection cells. Brine is stirred by convection, and its temperature homogenizes. Numerical computations (²Karimi-Jafari 2007) using Boussinesq' linearization of the equations describing natural convection in the cavern were performed. Axisymmetry is assumed. The ADINA CFD modulus was used (k-ε model, no slip at cavern walls, velocity profile at cavern wall according to the Reichard law). On Figure 2, steady state flow is reached (after $1.1 \cdot 10^8$ s, or 12 days). Iso-values of the stream function $\psi = \psi(r, \theta)$, $v_r = -\partial \psi / \partial \theta$, $v_\theta = \partial \psi / \partial r$ are represented on Figure 2, left: three convection cells develop, the lower cell rolls counter-clockwise. On Figure 2, right, velocity magnitude is represented; it is largest along the axis of symmetry where it is faster than 0.2 mm/s (17 m/day).

Figure 3 displays the case of a brine-filled cavern at Stassfurt, Germany. The cavern volume is $V_c = 13\,000 \text{ m}^3$, and the cavern's top and bottom are 440 m and 650 m, respectively. As in the previous example, the thermal gradient in the cavern is much smaller than the natural geothermal gradient, which is $G_R^\infty = 1.8 \times 10^{-2} \text{ }^\circ\text{C/m}$ in the salt formation — clear evidence of the onset of convection.

1.2.2. Gas-filled cavern

In a gas cavern, convection also can be observed. In fact, onset of convection is certain in a brine cavern (see Section 1.3). This is less true in a gas cavern and, in many cases, only the upper part of the cavern is the seat of perennial natural convection, as explained below.

Figure 4 represents the temperature profile in the S107 natural gas cavern at Stassfurt, Germany. ⁴Klafki et al. (2003) mention that the top and bottom of the cavern are 968 m and 1070 m, respectively. However, Fig. 4 suggests that actual depths might be slightly larger. Cavern volume is $V_c = 288\,000 \text{ m}^3$; it can be divided in three parts, whose radii are, from top to bottom, 20 m, 60-65 m and 46 m. The cavern had been kept at rest for 2.5 months before the temperature log was run. Thermal convection is active in the upper part of the cavern: the thermal gradient there is much smaller than the geothermal gradient in the rock mass ($G_R^\infty = 1.35 \times 10^{-2} \text{ }^\circ\text{C/m}$). In the lower part of the cavern, the gradient sign changes (gradient *inversion*) slightly above the cavern bottom. The lower part of the logging tool was below the

brine-gas interface when the temperature at the lowest point on Fig. 4 was measured, which means that this point is representative of the temperature at the interface depth. Gradient inversion will be discussed in Section 4.

1.3. First condition for onset of natural convection

Gas thermal constants are given in Table 1 (at $\theta = 25^\circ\text{C}$ and $P_N = 0.10325$ MPa). Heat capacity, C_p , is an increasing function of pressure. At 20 MPa and 30 °C, it is 3400 J/kg/°C for CH4 and 1600 J/kg/°C for air (Schlichtenmayer et al. 2015). In this paper, we are mainly interested in a semi-quantitative description of the phenomena observed in actual caverns – rather than in accurate numerical computations. For this reason, dependency of the thermal parameters upon temperature is not taken into account. Note, however, that it can play a role. For instance, it is known that, for a given geometry, rotation of convection cells is in opposite directions when a gas and a liquid is considered, as the sign of $d\mu/dT$ (μ is the fluid viscosity) is not the same.

Mechanical equilibrium in a fluid requires that temperature be a function of depth only. At equilibrium, $\text{grad } P = \rho(P,T)\vec{g}$; hence, $\text{curl}(\rho\vec{g}) = \left(\frac{\partial\rho}{\partial T}\right)\vec{g} \wedge \text{grad } T = \vec{0}$ and $T = T(z)$. Such an equilibrium is stable when the temperature gradient is oriented upward ($G = dT/dz < 0$). A downward-oriented temperature gradient ($G = dT/dz > 0$) is not a sufficient condition for onset of natural convection (i.e., instability). Even in an infinite fluid medium (No viscous boundary layer at cavern wall or viscous dissipation in the fluid body is considered.), this gradient must be larger than a certain threshold, the adiabatic gradient, which can be defined as follows. When a fluid particle rises by $dz < 0$, its pressure immediately decreases by $dP = \rho g dz$ to reach mechanical equilibrium with the surrounding fluid. Because of this expansion, liquid temperature drops. As a consequence, heat transfer from the surrounding fluid takes place to reach thermal equilibrium. However, this process is much slower than the mechanical process; in fact, when particle rise is fast, insufficient time is given for heat transfer, and particle expansion is adiabatic. In the case of a non-viscous ideal gas, an adiabatic expansion can be described as $Tds = C_p dT_{ad} - dP/\rho = 0$, where s is the gas entropy per unit of mass, C_p is the gas thermal capacity (when gas pressure is kept

constant), $\rho = P / rT$ is its density, from which $dT_{ad} = gdz / C_p$. This temperature change must be compared to the geothermal temperature change, or $dT_c = G_R^\infty dz$. When $G_{ad} = g / C_p < G_R^\infty$, the rising particle is warmer (and lighter) than its environment and keeps rising: convection appears. Gravity acceleration is $g = 10 \text{ m/s}^2$; heat capacity is $C_p = 3400 \text{ J/kg/}^\circ\text{C}$ for natural gas (at 30°C and 20 MPa), $C_p = 1100 \text{ J/kg/}^\circ\text{C}$ for air (at 30°C and 7 MPa), and $C_p = 14\,000 \text{ J/kg/}^\circ\text{C}$ for hydrogen. In other words, as $G_R^\infty = 1.6 - 1.8 \times 10^{-2} \text{ }^\circ\text{C/m}$ in a salt formation, onset of convection is likely in an air-storage facility ($g / C_p = 9 \times 10^{-3} \text{ }^\circ\text{C/m}$), highly likely in a natural gas storage ($g / C_p = 3 \times 10^{-3} \text{ }^\circ\text{C/m}$) and certain in a hydrogen storage ($g / C_p = 0.7 \times 10^{-3} \text{ }^\circ\text{C/m}$). In the case of a liquid, $dT_{ad} = \alpha T dP / \rho C_p$; $\alpha = -\partial \rho(T, P) / \partial T$ is the thermal expansion coefficient (for a gas, $\alpha T = 1$.) For brine, $\alpha_b = 4.4 \times 10^{-4} / ^\circ\text{C}$; $T = 300 \text{ K}$ is typical, $C_p^b = 3800 \text{ J/kg/}^\circ\text{C}$, and $G_{ad} = 3.5 \times 10^{-4} \text{ }^\circ\text{C/m}$ is much smaller than the geothermal gradient: onset of convection is certain in a brine-filled cavern, as suggested by the examples shown on Figs. 1 and 3. However, these simple arguments fail to take into account several factors, which will be discussed below.

1.4. Induced geothermal gradient

In the previous section, it was assumed that the geothermal gradient is the same in the cavern (G_c^∞) as in the rock mass (G_R^∞). In fact, because gas (or liquid) conductivity is less than rock conductivity, the geothermal gradient at rest (when no convection takes place) is larger in the cavern fluid than in the rock mass. The adiabatic gradient (G_{ad}) must be compared to G_c^∞ rather than G_R^∞ . Let T_R and T_F be the temperature distributions at rest in the rock mass and in the cavern fluid, respectively. Both are harmonic functions, $\Delta T_R = 0$ and $\Delta T_F = 0$. Consider, for instance, the case of a spherical cavern, radius a , whose origin is at the center of the cavern whose depth is H ; spherical coordinates are used, and $z = H + r \cos \theta$ is oriented downward. Rock temperature and fluid temperature are $T_R(r, \theta) = T_\infty(H) + G_R^\infty r \cos \theta + A a^3 \cos \theta / r^2$ and $T_F = T_\infty(H) + G_c^\infty r \cos \theta$, respectively; A and G_c^∞ are two constants. At the cavern wall, temperatures and heat flux are continuous, $T_R = T_F$ and $K_F \partial T_F / \partial n = K_R \partial T_R / \partial n$, from which it is inferred that the geothermal gradient at rest in the cavern is $G_c^\infty / G_R^\infty = 3K_R / (2K_R + K_F)$. In most cases, K_F is much smaller than K_R (for rock

salt, $K_R = 6 \text{ W/m/}^\circ\text{C}$; for brine, $K_F = 0.6 \text{ W/m/}^\circ\text{C}$; for air or for hydrogen, $K_F = 0.026 \text{ W/m/}^\circ\text{C}$ or $0.185 \text{ W/m/}^\circ\text{C}$ respectively — see Table 1) and the “induced” geothermal gradient in the fluid at rest is larger than the geothermal gradient in the rock mass (for air, $G_c^\infty = 1.5 G_R^\infty$), making onset of convection somewhat easier than predicted in Section 1.3. The induced geothermal gradient also depends on cavern age. During the leaching period, which is several years long, cold water pumped from a lake or a shallow aquifer is injected in the cavern. When leaching is completed, the cavern brine is colder than the rock mass by a couple of dozen of $^\circ\text{C}$. After debrining, the cavern walls are cold (⁷Walden et al., 2001), and the same can be said of the brine left at the cavern bottom. Warming brine and rock mass is slow (⁸Fosse and Røvang, 1998), as discussed later. The transient induced geothermal gradient is difficult to compute.

1.5. Wet gas and dry gas

In many cases (Section 2), cavern gas is saturated with water vapor. (Equilibrium with sump brine must be reached.) In principle, when computing adiabatic gas expansion, the effect of condensation must be taken into account, and the wet adiabatic gradient is smaller than the dry adiabatic gradient (⁹Sivoukhine, 1982, p. 499-501). Let m_g , m_v , m_L be the mass of gas, water vapor and liquid water, respectively. When condensation starts, $m_L = 0$ and $\dot{m}_L + \dot{m}_v = 0$. The vapor pressure at saturation, $P_v^s(T)$, is a function of temperature only. During adiabatic evolution, the entropy of the mixture, $s = \sum_i m_i s_i$, $i = v, L, g$, is constant, and

$$(s_v - s_L)dm_v + \frac{m_v}{T}(C_P^v dT_{ad} - \frac{dP_v^s}{\rho_v}) + \frac{m_g}{T}(C_P^g dT_{ad} - \frac{dP_g}{\rho_g}) = 0 \quad (1)$$

Let $L = (s_v - s_L)T$ be the latent heat capacity for the liquid-vapor transition; taking into account $dP_v^s / dT = L \rho_v / T$ (Capeyron’s law), $dP_g = \rho_g g dz$ (mechanical equilibrium), and $P_i = \rho_i r_i T$ for $i = v, g$ (state equations)

$$\left. \frac{dT}{dz} \right|_{wet} = \frac{g}{C_p^g \left(1 + \frac{r_g P_v^s}{r_v P_g} \left[L \left(\frac{L}{r_v T^2} - \frac{2}{T} \right) + C_p^v \right] \right) / C_p^g} \quad (2)$$

The difference between the wet adiabatic gradient G_{ad}^{wet} and the dry adiabatic gradient G_{ad}^{dry} (which equals g / C_p^v) is proportional to the P_v^s / P_g ratio. This ratio between vapor pressure and gas pressure is not very small in the atmosphere ($P_g = P_{air} = 0.1$ MPa; at $\theta = 30^\circ\text{C}$, $P_v^s(30^\circ\text{C}) = 4.25$ kPa and $P_v^s / P_{air} = 4\%$, see Section 2) and the difference is significant when meteorological phenomena are concerned. The same is not true for a gas cavern, in which $P_g = 10$ MPa, and, when $\theta = 50^\circ\text{C}$, $P_v^s(50^\circ\text{C}) = 9$ kPa are typical.

1.6. Second condition for onset of natural convection

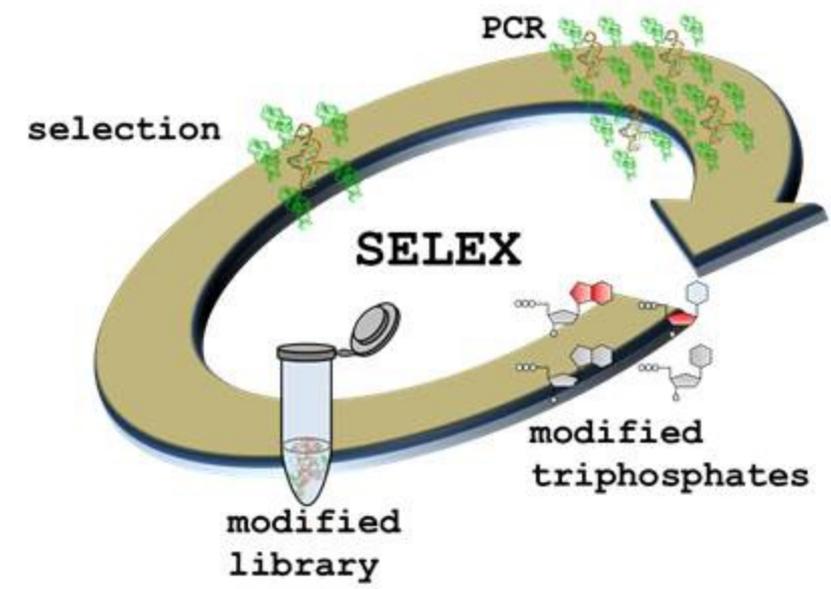
The condition for onset of natural convection, $G_{ad} < G_c^\infty$, is a necessary condition. In the analysis, viscous effects were not taken into account. In fact, in a cavity of (very) small dimensions, viscosity can impede natural convection. The set of equations describing natural convection in a cavern (conservations of mass, energy and momentum) must be solved numerically. However, steady-state convective flow can be described through linearized equations (Boussinesq approximation) (¹⁰Landau and Lifschitz, 1971). This set of equations highlights the significance of three dimensionless constants, Prandtl, Grashof and Nusselt numbers: $Pr = \nu/k$, where ν = kinematic viscosity, k = gas thermal diffusivity; $Gr = g \alpha G a^4 / \nu^2$, where a = cavern characteristic length, α = the thermal expansion coefficient of the fluid, and $G = G_c - G_{ad}$, where G_c is the actual thermal gradient in the cavern; and $Nu = \bar{H} a / K_f$ where \bar{H} is the heat-transfer coefficient at the cavern wall (defined in Section 5.1). Natural convection starts when G is larger than the value given by a certain combination of Gr, Pr and Nu that depends on cavern shape and boundary conditions. For an elongated cylindrical cavern, convection appears when the Rayleigh number $Ra = Pr \cdot Gr$ is larger than a value that depends on Nu and is smaller than 10^3 (¹⁰Landau and Lifschitz, 1971). The Prandtl number can be written $Pr = \mu C_p / K$; for most gases (Table 1), dynamic viscosity is in the range $\mu = 1 - 2 \times 10^{-5}$ Pa.s $C_p = 1 - 2 \times 10^3$ J/kg/K, and $K = 2 \times 10^{-5}$ Pa.s. Also, $g = 10$ m/s², $G = 1.5 \times 10^{-2}$ °C/m, $\alpha_g = 3 \times 10^{-3}$ /°C and, for standard caverns, $a = 10$ -50 m.

The Rayleigh number ($=Pr.Gr$) is larger than 10^{10} , a figure for which natural convection is turbulent. In other words, a cavern contains so large a fluid body that the condition relative to the adiabatic gradient is a sufficient condition for onset of convection. For instance, ⁴Klafki et al., 2003 computed natural convection in the cavern represented on Fig. 4. A stable vortex (The horizontal component of gas speed was high.) was observed in the lower part of the cavern, an independent convection cell (upstream at the wall, downstream in the center) was observed in the main cavern, and flow was unsteady in the narrow upper part. Maximal velocities were smaller than 0.5 m/s. This computation highlights the importance of cavern shape on the pattern of convection cells pattern (see also Fig. 2).

2. Water vapor in the cavern

Moisture content is a critical issue when operating a gas cavern. It was proven in Section 1.5 that its influence on the onset of natural convection is minor. However, when gas containing water vapor is withdrawn from a cavern, gas pressure and temperature drop in the wellbore, leading to the possible formation of hydrates in the upper part of the wellbore (a severe operational problem, discussed in Section 6). In fact, natural gas injected in a cavern is relatively dry. In France, for instance, water vapor concentration (or $c_v = m/V_N = P_v P_N / r_v T_N P_g$, where the subscript N is for normal conditions, $P_N = 0.10325$ MPa, and $T_N = 273$ K.) must be less than $c_v = 46$ mg/Nm³ (¹¹Louvet et al., 2018). However, there is a thin brine layer at the cavern wall (70 cm³/m², see ¹²Kökritz and Walden, 1994) and a (much larger) amount of brine in the sump at the cavern bottom mentioned in Section 1.1. Equilibrium requires, at least in principle, that brine is saturated with dissolved natural gas and that cavern gas is saturated with water vapor. The partial pressure of water vapor at saturation, or P_v^s , is a function of temperature (for instance, Rankine's formula predicts: $P_v^s [T] = f P_{am} \exp(13.7 - 5120/T)$, with T in K; for pure water, $f = 1$, and for saturated brine, $f = 0.75$, according to Raoult's law).

Following gas injection or withdrawal, temperature changes in the cavern, and gas is either over- or under-saturated with water vapor in the new temperature conditions. Water vaporizes or condenses until, after some time, equilibrium is reached again. Consider an air-storage cavern of volume $V_c = 400\,000$ m³: air is withdrawn rapidly from the cavern, whose



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